

## Key words you need to know

Relative isotopic mass, relative atomic mass, isotope, proton, neutron, electron, atom, atomic structure

## Find out about

Describe the process by which new theories are accepted by scientists.

## Do you remember?

How has the model of the atom changed over the years and still continues to do so?

How do you deduce the atomic structure in atoms and ions?

## Practice Practice Practice

Draw a table to describe, the charge, position and mass of protons, neutrons and electrons. Include how these are calculated.

Explain, using an example, the term isotope, be sure to include the atomic structure.

Explain why  $^{12}\text{C}$  is used as the standard measurement of relative masses.

Explain, using an example how you calculate relative atomic masses. (\*Show ALL workings.)

Explain, with examples, how you work out relative molecular masses and relative formula masses. (\*Show ALL workings.)

## Exam questions . . .

Foundation Chemistry 2811:

- Jan 2001 1 a, b, c
- June 2001 4a
- May 2002 1a, b, c

## Extend your knowledge!

Explain how and why our scientific knowledge about the atoms is evolving.

